

## **Alkalinity**

Alkalinity is related to pH. It is a measure of the water's ability to provide a stable pH level and to avoid rapid changes in pH that could adversely affect the health of the aquatic life. Fish exposed to changes in pH outside their normal range can be stressed or even die. Stress leaves fish vulnerable to disease, degrading their health. Very rapid changes in pH can cause fish to lose control over their swim bladders, making it hard for them to swim correctly. Additionally, alkaline conditions can transform nitrogen in the water column into a more toxic form of ammonia that can interfere with a fish's ability to breathe normally and at high concentrations can poison fish. Withdrawals from the stream will reduce the stream's heat capacity and cause greater fluctuation in daytime and nighttime stream temperatures. When substantial plant or algal growth are present, this will lead to greater fluctuations in alkalinity and pH. Additional withdrawals from a stream that is already impaired for pH will exacerbate these problems. Fish and aquatic insects are sensitive to imbalances in pH. Low pH levels (below 5) may lead to death and high pH levels (9-14) can harm fish by denaturing cellular membranes. These alkalinity and pH imbalances will result in the diminution of the habitat of sensitive, threatened, or endangered fish species.

## **Ammonia**

Ammonia is a toxic form of nitrogen. Commercial fertilizers and decomposing organic wastes are common sources of ammonia that enter water bodies through rainfall and runoff. Toxicity increases with higher pH and higher temperatures. In water bodies with high levels of nutrients and algae, respiration and photosynthesis cause extreme diurnal swings in pH. When pH is highest during the day, ammonia toxicity is high. Ammonia enters the blood and tissues of fish and aquatic life and disrupts cell functions. When ammonia toxicity is high, fish and aquatic life experience reduced feeding, lowered reproduction, disrupted organ function, and even death.

## **Aquatic weeds and/or algae**

Both rooted aquatic plants and algae are a natural part of stream systems. They grow by taking in nutrients from the water column and sunlight. When water temperatures are warm enough and sufficient nutrients are present, excessive growth can occur; this can be a problem for both aquatic life and recreational beneficial uses. Excessive growth can affect aquatic life in several ways. During sunlight hours, plants and algae remove carbon dioxide from the water column as part of photosynthesis. With excessive growth, this can result in increased pH (alkaline conditions). During the night, plant growth removes oxygen from water and releases carbon dioxide, resulting in both low pH (acidic conditions) and low dissolved oxygen. In addition, when algae die and decompose they remove oxygen from the surrounding water. Low dissolved oxygen can lead to decreased fish habitat and even fish kills. Additionally, low dissolved oxygen levels can lead to changes in water chemistry that allow mercury to be more able to enter the food chain. Algal blooms also often create odors and coloration that are objectionable to recreational users. A reduction in stream flow would result in increased water temperature and increased nutrient concentrations, both of which would contribute to a greater risk of excessive plant growth and algal blooms. Reduced stream flow would also result in reduced flushing capacity (to remove decomposing plant and algal materials) which would exacerbate conditions in following years. Additionally, decreased stream flow would increase the occurrence of low dissolved

oxygen from plant growth and decomposition and increase the opportunity for mercury to enter the food chain.

### **Biological Criteria**

Oregon's biological criteria standards are based on the assemblage of species needed to maintain a healthy resident biological community. Resident biological communities are the local food webs that support fish. Reduced flows and increased temperatures will degrade the biological community and therefore result in the diminution of habitat of sensitive, threatened, or endangered fish species.

### **Dissolved Oxygen**

Fish and other aquatic organisms require different concentrations of dissolved oxygen based on their species and life history stage. Oregon's dissolved oxygen standards are based on the most sensitive species and life history stage at the location and season of concern. Dissolved oxygen levels are affected by temperature, flow, nutrient loading, algae growth, and other factors. If dissolved oxygen drops too low enough levels, it can result in fish kills. In waterbodies where dissolved oxygen concentrations are known to be insufficient for the habitat of sensitive, threatened, and endangered fish, any additional reduction in dissolved oxygen concentrations would result in the diminution of habitat.

### **E. Coli**

E. coli is a bacteria that lives in the intestines of humans and animals. E. coli enters water bodies from agricultural runoff, stormwater runoff, failing septic systems, and wildlife. Humans and animals can get sick from harmful types of E. coli by ingesting water while swimming or recreating in the water body. Higher concentrations of E. coli are linked to higher concentrations of harmful bacteria and viruses in the water body. High concentrations of E. coli and associated pathogens result in more illnesses, and more severe illnesses, in humans and animals in contact with the water body.

### **Iron**

Iron is common in many rocks and is an important component of many soils. Iron is an essential trace element required by both plants and animals. Ferrous ( $\text{Fe}^{2+}$ ) and ferric ( $\text{Fe}^{3+}$ ) irons are the primary forms of concern in the aquatic environment. Ferrous iron is colorless (clear) while ferric iron will show up as a rust colored stain in the water. Iron bacteria may also be present in streams associated with mining waste or ground water recharge. A rust-colored slime often forms rocks and other surfaces when iron bacteria are present. Iron and manganese often occur together. High concentrations of these metals can result in discolored water. Where water supplies are used for domestic purposes, elevated iron and manganese concentrations can result in stained plumbing fixtures and an unpleasant metallic taste to the water. Iron deposits can buildup in pressure tanks, storage tanks, water heaters, and pipelines, decreasing capacity, reducing pressure, and increasing maintenance. Iron and manganese concentrations of concern are generally established on the basis of aesthetic and economic considerations (unpleasant tastes and coloration) rather than toxicity. A reduction in streamflow will lead to an increased concentration of iron and manganese in the water column. This may result in increased bacterial growth and an increase in aesthetic, recreational and domestic water system impacts.

## **Manganese**

Manganese is a metal found naturally in rocks and soil. It does not occur as a pure element in nature, but always combines with oxygen or other elements. Manganese is an essential trace element required by both plants and animals. Iron and manganese often occur together. High concentrations of these metals can result in discolored water. Where water supplies are used for domestic purposes, elevated iron and manganese concentrations can result in stained plumbing fixtures and an unpleasant metallic taste to the water. Iron deposits can buildup in pressure tanks, storage tanks, water heaters, and pipelines, decreasing capacity, reducing pressure, and increasing maintenance. Iron and manganese concentrations of concern are generally established on the basis of aesthetic and economic considerations (unpleasant tastes and coloration) rather than toxicity. A reduction in streamflow will lead to an increased concentration of iron and manganese in the water column. This may result in increased bacterial growth and an increase in aesthetic, recreational and domestic water system impacts.

## **pH**

pH is a measure of how acidic or basic (alkaline) the water is. Water with a pH greater than 7 is alkaline, water with a pH of less than 7 is acidic. Every species of fish has adapted to a specific range of pH. Fish exposed to changes in pH outside their normal range can be stressed or even die. Stress leaves fish vulnerable to disease, degrading their health. Additionally, alkaline conditions can transform nitrogen in the water column into a more toxic form of ammonia that can poison fish. Withdrawals from the stream will reduce the stream's heat capacity and cause greater fluctuation in daytime and nighttime stream temperatures. When nutrients and sunlight are sufficiently present, higher stream temperatures lead to more algal growth. During the day, algae absorb carbon dioxide from the water for cell growth, raising pH. At night, photosynthesis stops and algae continue to respire, releasing carbon dioxide and lowering pH. This cycle creates diel fluctuations in pH. Additional withdrawals from a stream that is already impaired for pH will lead to larger diel fluctuations in pH. Fish and aquatic insects are sensitive to imbalances in pH. Low pH levels (below 5) may lead to death and high pH levels (9-14) can harm fish by denaturing cellular membranes. These pH imbalances result in the diminution of the habitat of sensitive, threatened, or endangered fish species.

## **Phosphorus, Phosphate Phosphorus**

Phosphorus is an essential plant nutrient, but an excess of phosphorus can be detrimental to aquatic life. High phosphorus concentrations can lead to eutrophication, a situation where aquatic plants grow so rapidly that dissolved oxygen concentrations drop below the levels needed to sustain fish and other aquatic life. Phosphate (also referred to as orthophosphate) is a chemical form of phosphorus that is very soluble and readily available for plant uptake, leading to rapid growth and, in the case of algae, rapid expansion of algal blooms. A reduction in streamflow will increase phosphorus concentrations. This would cause longer or more severe instances of oxygen depletion, resulting in a diminution of water quality for the habitat of sensitive, threatened, or endangered fish species.

## **Sedimentation**

While sediment is an essential part of healthy functioning stream systems, excessive sediment loads can

have severe negative impacts on a stream ecosystem. Many fish species are adapted to high suspended sediment levels that occur for short periods of time, but longer exposure to high levels of suspended sediment can interfere with feeding behavior, damage gills, reduce available food, and reduce growth rates. Deposition and sedimentation (when sediment falls out of the water column and deposits on the streambed) can smother eggs and fry in the substrate and fill in pools within the stream channel (reducing or eliminating cold water refugia important to cold water aquatic life during periods of high water temperature). Because bacteria, nutrients and other chemical substances are often attached to sediment particles, excessive sediment loading can also increase nutrient and toxic concentrations and contribute to decreased dissolved oxygen in both the water column and the spawning gravels. A reduction in streamflow will lead to locally increased deposition and sedimentation. It will also result in an increased rate of evaporation in warm weather, which in turn can increase nutrient and toxic concentrations in the stream. This would result in the diminution of water quality for the habitat of sensitive, threatened, or endangered fish species.

### **Temperature**

Oregon's stream temperature standards are based on the life cycle needs of salmonids. Stream temperatures that exceed the standards can disrupt the life cycle of a sensitive, threatened, or endangered fish species and may even cause death. Further withdrawals from the stream will reduce the stream's heat capacity and cause greater fluctuation in daytime and nighttime stream temperatures. This will result in the diminution of habitat of sensitive, threatened, or endangered fish species.